

| Reacción | Constante | Temperatura °C |
|--|--|----------------|
| $\text{EtOH} + \text{HAc} \rightleftharpoons \text{AcEt} + \text{H}_2\text{O}$ | $K_C=4$ | 25 |
| $2\text{NO}_{2(g)} \rightleftharpoons 2\text{NO}_{(g)} + \text{O}_{2(g)}$ | $K_C=1,42 \cdot 10^{-4}$ $K_P=6,98 \cdot 10^{-3}$ | 327 327 |
| $\text{N}_2\text{O}_{4(g)} \rightleftharpoons 2\text{NO}_{2(g)}$ | $K_C=0,142$ $K_P=0,32$ | 25 35 |
| $\text{H}_{2(g)} + \text{I}_{2(g)} \rightleftharpoons 2\text{HI}_{(g)}$ | $K_C=56,6$ | 450 |
| $2\text{HI}_{(g)} \rightleftharpoons \text{H}_{2(g)} + \text{I}_{2(g)}$ | $K_C=0,019$ $K_C=0,022$ | 350 490 |
| $\text{PCl}_{5(g)} \rightleftharpoons \text{PCl}_{3(g)} + \text{Cl}_{2(g)}$ | $K_C=0,050$ $K_P=2,144$ | 250 250 |
| $\text{N}_2\text{O}_4 \rightleftharpoons 2\text{NO}_2$ (Dis. de cloroformo) | $K_C=1,07 \cdot 10^{-5}$ | 25 |
| $\text{N}_{2(g)} + \text{O}_{2(g)} \rightleftharpoons 2\text{NO}_{(g)}$ | $K_C=0,0120$ $K_P=5,33 \cdot 10^{-31}$ | 2800 25 |
| $\text{C}_{(s)} + \text{CO}_{2(g)} \rightleftharpoons 2\text{CO}_{(g)}$ | $K_P=122$ | 1000 |
| $\text{CO}_{2(g)} + \text{H}_{2(g)} \rightleftharpoons \text{CO}_{(g)} + \text{H}_2\text{O}_{(g)}$ | $K_P=1,59$ | 1000 |
| $\text{C}_{(s)} + \text{H}_2\text{O}_{(g)} \rightleftharpoons \text{CO}_{(g)} + \text{H}_{2(g)}$ | $K_P=76,4$ | 1000 |
| $2\text{SO}_{2(g)} + \text{O}_{2(g)} \rightleftharpoons 2\text{SO}_{3(g)}$ | $K_C=279$ $K_C=729$ | 727 550 |
| $\text{CaCO}_{3(s)} \rightleftharpoons \text{CaO}_{(s)} + \text{CO}_{2(g)}$ | $K_P=0,25$ $K_C=2,84 \cdot 10^{-3}$ | 800 800 |
| | $K_P=7,08 \cdot 10^{-4}$ | 327 |
| | $K_P=4,30 \cdot 10^{-5}$ | 427 |
| $\text{N}_{2(g)} + 3\text{H}_{2(g)} \rightleftharpoons 2\text{NH}_{3(g)}$ | $K_P=6,64 \cdot 10^5$ $K_C=100$ | 25 257 |
| | $K_P=2,97 \cdot 10^{-6}$ | 1000 |
| $2\text{H}_2\text{O}_{(g)} \rightleftharpoons 2\text{H}_{2(g)} + \text{O}_{2(g)}$ | $K_C=9,3 \cdot 10^{-12}$ | 1000 |
| $2\text{HCl}_{(g)} \rightleftharpoons \text{Cl}_{2(g)} + \text{H}_{2(g)}$ | $K_C=10^{-7}$ | 1000 |
| $4\text{HCl}_{(g)} + \text{O}_{2(g)} \rightleftharpoons 2\text{Cl}_{2(g)} + 2\text{H}_2\text{O}_{(g)}$ | $K_C=1,1 \cdot 10^{-3}$ | 1000 |
| $\text{CO}_{(g)} + \text{H}_2\text{O}_{(g)} \rightleftharpoons \text{CO}_{2(g)} + \text{H}_{2(g)}$ | $K_P=5$ | 530 |
| $\text{H}_{2(g)} + \text{Cl}_{2(g)} \rightleftharpoons 2\text{HCl}_{(g)}$ | $K_C=3,17 \cdot 10^{16}$ | 27 |
| $\text{COCl}_{2(g)} \rightleftharpoons \text{CO}_{(g)} + \text{Cl}_{2(g)}$ | $K_P=0,3178$ $K_C=4,77 \cdot 10^{-3}$ | 538 538 |