

Reacción	Constante	Temperatura °C
$\text{EtOH} + \text{HAc} \rightleftharpoons \text{AcEt} + \text{H}_2\text{O}$	$K_C=4$	25
$2\text{NO}_{2(g)} \rightleftharpoons 2\text{NO}_{(g)} + \text{O}_{2(g)}$	$K_C=1,42 \cdot 10^{-4}$	327
	$K_P=6,98 \cdot 10^{-3}$	327
$\text{N}_2\text{O}_{4(g)} \rightleftharpoons 2\text{NO}_{2(g)}$	$K_C=0,142$	25
	$K_P=0,32$	35
$\text{H}_{2(g)} + \text{I}_{2(g)} \rightleftharpoons 2\text{HI}_{(g)}$	$K_C=56,6$	450
$2\text{HI}_{(g)} \rightleftharpoons \text{H}_{2(g)} + \text{I}_{2(g)}$	$K_C=0,019$	350
	$K_C=0,022$	490
$\text{PCl}_{5(g)} \rightleftharpoons \text{PCl}_{3(g)} + \text{Cl}_{2(g)}$	$K_C=0,050$	250
	$K_P=2,144$	250
$\text{N}_2\text{O}_4 \rightleftharpoons 2\text{NO}_2$ (Dis. de cloroformo)	$K_C=1,07 \cdot 10^{-5}$	25
$\text{N}_{2(g)} + \text{O}_{2(g)} \rightleftharpoons 2\text{NO}_{(g)}$	$K_C=0,0120$	2800
	$K_P=5,33 \cdot 10^{-31}$	25
$\text{C}_{(s)} + \text{CO}_{2(g)} \rightleftharpoons 2\text{CO}_{(g)}$	$K_P=122$	1000
$\text{CO}_{2(g)} + \text{H}_{2(g)} \rightleftharpoons \text{CO}_{(g)} + \text{H}_2\text{O}_{(g)}$	$K_P=1,59$	1000
$\text{C}_{(s)} + \text{H}_2\text{O}_{(g)} \rightleftharpoons \text{CO}_{(g)} + \text{H}_{2(g)}$	$K_P=76,4$	1000
$2\text{SO}_{2(g)} + \text{O}_{2(g)} \rightleftharpoons 2\text{SO}_{3(g)}$	$K_C=279$	727
	$K_C=729$	550
$\text{CaCO}_{3(s)} \rightleftharpoons \text{CaO}_{(s)} + \text{CO}_{2(g)}$	$K_P=0,25$	800
	$K_C=2,84 \cdot 10^{-3}$	800
	$K_P=7,08 \cdot 10^{-4}$	327
$\text{N}_{2(g)} + 3\text{H}_{2(g)} \rightleftharpoons 2\text{NH}_{3(g)}$	$K_P=4,30 \cdot 10^{-5}$	427
	$K_P=6,64 \cdot 10^5$	25
	$K_C=100$	257
	$K_P=2,97 \cdot 10^{-6}$	1000
$2\text{H}_2\text{O}_{(g)} \rightleftharpoons 2\text{H}_{2(g)} + \text{O}_{2(g)}$	$K_C=9,3 \cdot 10^{-12}$	1000
$2\text{HCl}_{(g)} \rightleftharpoons \text{Cl}_{2(g)} + \text{H}_{2(g)}$	$K_C=10^{-7}$	1000
$4\text{HCl}_{(g)} + \text{O}_{2(g)} \rightleftharpoons 2\text{Cl}_{2(g)} + 2\text{H}_2\text{O}_{(g)}$	$K_C=1,1 \cdot 10^{-3}$	1000
$\text{CO}_{(g)} + \text{H}_2\text{O}_{(g)} \rightleftharpoons \text{CO}_{2(g)} + \text{H}_{2(g)}$	$K_P=5$	530
$\text{H}_{2(g)} + \text{Cl}_{2(g)} \rightleftharpoons 2\text{HCl}_{(g)}$	$K_C=3,17 \cdot 10^{16}$	27
$\text{COCl}_{2(g)} \rightleftharpoons \text{CO}_{(g)} + \text{Cl}_{2(g)}$	$K_P=0,3178$	538
	$K_C=4,77 \cdot 10^{-3}$	538