

Reacción	Constante	Temperatura
$\text{EtOH} + \text{HAc} \rightleftharpoons \text{AcEt} + \text{H}_2\text{O}$	$K_C=4$	25 °C
$2\text{NO}_{2(g)} \rightleftharpoons 2\text{NO}_{(g)} + \text{O}_{2(g)}$	$K_C=1,42 \cdot 10^{-4}$ $K_P=6,98 \cdot 10^{-3}$	327 °C 327 °C
$\text{N}_2\text{O}_{4(g)} \rightleftharpoons 2\text{NO}_{2(g)}$	$K_C=0,142$ $K_P=0,32$	25 °C 35 °C
$\text{H}_{2(g)} + \text{I}_{2(g)} \rightleftharpoons 2\text{HI}_{(g)}$	$K_C=56,6$	450 °C
$2\text{HI}_{(g)} \rightleftharpoons \text{H}_{2(g)} + \text{I}_{2(g)}$	$K_C=0,019$ $K_C=0,022$	350 °C 490 °C
$\text{PCl}_{5(g)} \rightleftharpoons \text{PCl}_{3(g)} + \text{Cl}_{2(g)}$	$K_C=0,050$ $K_P=2,144$	250 °C 250 °C
$\text{N}_2\text{O}_4 \rightleftharpoons 2\text{NO}_2$ (Dis. de cloroformo)	$K_C=1,07 \cdot 10^{-5}$	25 °C
$\text{N}_{2(g)} + \text{O}_{2(g)} \rightleftharpoons 2\text{NO}_{(g)}$	$K_C=0,0120$ $K_P=5,33 \cdot 10^{-31}$	2800 °C 25 °C
$\text{C}_{(s)} + \text{CO}_{2(g)} \rightleftharpoons 2\text{CO}_{(g)}$	$K_P=122$	1000 °C
$\text{CO}_{2(g)} + \text{H}_{2(g)} \rightleftharpoons \text{CO}_{(g)} + \text{H}_2\text{O}_{(g)}$	$K_P=1,59$	1000 °C
$\text{C}_{(s)} + \text{H}_2\text{O}_{(g)} \rightleftharpoons \text{CO}_{(g)} + \text{H}_{2(g)}$	$K_P=76,4$	1000 °C
$2\text{SO}_{2(g)} + \text{O}_{2(g)} \rightleftharpoons 2\text{SO}_{3(g)}$	$K_C=279$ $K_C=729$	727 °C 550 °C
$\text{CaCO}_{3(s)} \rightleftharpoons \text{CaO}_{(s)} + \text{CO}_{2(g)}$	$K_P=0,25$ $K_C=2,84 \cdot 10^{-3}$	800 °C 800 °C
$\text{N}_{2(g)} + 3\text{H}_{2(g)} \rightleftharpoons 2\text{NH}_{3(g)}$	$K_P=7,08 \cdot 10^{-4}$ $K_P=4,30 \cdot 10^{-5}$ $K_P=6,64 \cdot 10^5$ $K_C=100$ $K_P=2,97 \cdot 10^{-6}$	327 °C 427 °C 25 °C 257 °C 1000 °C
$2\text{H}_2\text{O}_{(g)} \rightleftharpoons 2\text{H}_{2(g)} + \text{O}_{2(g)}$	$K_C=9,3 \cdot 10^{-12}$	1000°C
$2\text{HCl}_{(g)} \rightleftharpoons \text{Cl}_{2(g)} + \text{H}_{2(g)}$	$K_C=10^{-7}$	1000 °C
$4\text{HCl}_{(g)} + \text{O}_{2(g)} \rightleftharpoons 2\text{Cl}_{2(g)} + 2\text{H}_2\text{O}_{(g)}$	$K_C=1,1 \cdot 10^{-3}$	1000 °C
$\text{CO}_{(g)} + \text{H}_2\text{O}_{(g)} \rightleftharpoons \text{CO}_{2(g)} + \text{H}_{2(g)}$	$K_P=5$	530 °C
$\text{H}_{2(g)} + \text{Cl}_{2(g)} \rightleftharpoons 2\text{HCl}_{(g)}$	$K_C=3,17 \cdot 10^{16}$	27 °C
$\text{COCl}_{2(g)} \rightleftharpoons \text{CO}_{(g)} + \text{Cl}_{2(g)}$	$K_P=0,3178$ $K_C=4,77 \cdot 10^{-3}$	538 °C 538 °C