Acid	Reaction	k_a
Perchloric	$HClO_4 + H_2O \rightarrow ClO_4^- + H_3O^+$	Very high
Hydroiodic	$HI + H_2O \rightarrow I^- + H_3O^+$	Very high
Hydrobromic	$HBr + H_2O \rightarrow Br^- + H_3O^+$	Very high
Hydrochloric	$HCl + H_2O \rightarrow Cl^- + H_3O^+$	Very high
Sulfuric	$H_2SO_4 + H_2O \rightarrow HSO_4^- + H_3O^+$	Very high
Nitric	$HNO_3 + H_2O \rightarrow NO_3^- + H_3O^+$	Very high
Iodic	$HIO_3 + H_2O \rightleftharpoons IO_3^- + H_3O^+$	0.19
Sulfurous	$H_2SO_3 + H_2O \rightleftharpoons HSO_3^- + H_3O^+$	0.016
Hydrogen sulfate	$HSO_4^- + H_2O \rightleftharpoons SO_4^{=} + H_3O^+$	0.012
Phosphoric	$H_3PO_4 + H_2O \rightleftharpoons H_2PO_4^- + H_3O^+$	$7.5 \cdot 10^{-3}$
Hydrofluoric	$HF + H_2O \rightleftharpoons F^- + H_3O^+$	$7.0 \cdot 10^{-4}$
Nitrous	$HNO_2 + H_2O \rightleftharpoons NO_2^- + H_3O^+$	$4.5 \cdot 10^{-4}$
Formic	$HCOOH + H_2O \rightleftharpoons HCOO^- + H_3O^+$	$1.8 \cdot 10^{-4}$
Benzoic	$C_6H_5COOH + H_2O \rightleftharpoons C_6H_5COO^- + H_3O^+$	$6.6 \cdot 10^{-5}$
Acetic	$CH_3COOH + H_2O \rightleftharpoons CH_3COO^- + H_3O^+$	$1.8 \cdot 10^{-5}$
Carbonic	$H_2CO_3 + H_2O \rightleftharpoons HCO_3^- + H_3O^+$	$4.2 \cdot 10^{-7}$
Hydrosulfuric	$H_2S + H_2O \rightleftharpoons HS^- + H_3O^+$	$1.0 \cdot 10^{-7}$
Hypochlorous	$HClO + H_2O \rightleftharpoons ClO^- + H_3O^+$	$3.0 \cdot 10^{-8}$
Dihydrogen phosphate	$H_2PO_4^- + H_2O \rightleftharpoons HPO_4^{=} + H_3O^+$	$6.2 \cdot 10^{-8}$
Ammonium	$NH_4^+ + H_2O \rightleftharpoons NH_3 + H_3O^+$	$5.5 \cdot 10^{-10}$
Hydrogen cyanide	$HCN + H_2O \rightleftharpoons CN^- + H_3O^+$	$4.0 \cdot 10^{-10}$
Hydrogen carbonate	$HCO_3^- + H_2O \rightleftharpoons CO_3^{=} + H_3O^+$	$5.6 \cdot 10^{-11}$
Hydrogen phosphate	$HPO_4^{=} + H_2O \rightleftharpoons PO_4^{3-} + H_3O^+$	$2.2 \cdot 10^{-13}$
Hydrogen sulfide	$HS^- + H_2O \rightleftharpoons S^{=} + H_3O^+$	$1.0 \cdot 10^{-13}$

Base	Reaction	k_b
Sodium hydroxide	$NaOH \rightarrow Na^+ + OH^-$	Very high
Potassium hydroxide	$KOH \rightarrow K^+ + OH^-$	Very high
Methylamine	$CH_3NH_2 + H_2O \rightleftharpoons CH_3NH_3^+ + OH^-$	$5.4 \cdot 10^{-4}$
Ammonia	$NH_3 + H_2O \rightleftharpoons NH_4^+ + OH^-$	$1.8 \cdot 10^{-5}$
Aniline	$C_6H_5NH_2 + H_2O \rightleftharpoons C_6H_5NH_3^+ + OH^-$	$3.8 \cdot 10^{-10}$